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Solution dilution worksheet

Chemistry WS 9-3 Dilutions Worksheet

Name: _____ Period: _____

$$M_1V_1 = M_2V_2$$

where:
 M_1 = molarity of concentrated stock sol'n
 M_2 = molarity of diluted sol'n
 V_1 = volume of the concentrated stock sol'n
 V_2 = volume of the diluted sol'n

- When a concentrated stock solution is diluted to prepare a less concentrated reagent, the number of _____ is the same before and after dilution.
- When the volume of a given solution is doubled by adding water the new concentration is the original concentration.
- Calculate the new molarity that results when 250.0mL of water is added to each of the following solutions (add the volume of water to the existing volume and divide moles by the new volume):
 - 125mL of .251M HCl (0.084M)
 - 445 mL of 0.499M H₂SO₄ (0.320M)
 - 5.25 mL of 0.101M HNO₃ (0.0021M)
 - 11.2 mL of 14.5M HC₂H₃O₂ (0.62M)
- Concentrated sulfuric acid is typically 18.1M H₂SO₄. Calculate the volume in mL of concentrated sulfuric acid needed to prepare 125mL of 0.100M H₂SO₄ solution. (0.691mL of 18.1M H₂SO₄)
- If you add 25 mL of water to 125 mL of a 0.15 M NaOH solution, what will the molarity of the diluted solution be? (0.125)

Solutions – Dilutions - Concentration Worksheet

Formula & answer statement for making dilutions From Liquids Formula and Answer Statement for making solutions from solids.

- Explain how to make a 5.85 L sample of 0.054 M Nitric acid. HNO₃
- Explain how to make a 350 ml sample of 3.00 M potassium nitrate solution. KNO₃
- Find the molarity if 5.125 moles of calcium iodide is dissolved in a 4.00 L solution. CaI₂
- How many grams of KNO₃ are needed to make 3.00 L of 5 M KNO₃?
- How many liters of 4M solution can be made using 100 grams of LiBr?
- What is the concentration of a solution that has a volume of 2.5 L and contains 660 grams of Ca₃(PO₄)₂?
- How much 0.075 M NaCl solution can be made by diluting 450 mL of 9.0M NaCl?
- If 550 mL of a 3.50 M KCl solution is set aside and allowed to evaporate until the volume of the solution is 275 mL, what will the molarity of the solution be?
- What is the molarity of a solution in which 0.45 grams of NaNO₃ are dissolved in 265mL of solution.
- If I add 25 mL of water to 125 mL of a 0.15 M NaOH solution, what is the molarity of the diluted solution?
- How much water would I need to add to 500 mL of a 2.4 M KCl solution to make a 1.0 M solution?
- Explain how 450 mL of a 0.250 M NaOH solution is made.
- Explain how 750 mL of a 0.100 M CaCl₂ solution is made?
- Explain how to make 125 mL of a 0.050 M BeCl₂ solution?
- Explain how to make 30.0 mL of 0.35 M HCl solution from 12 M HCl.
- Explain how to make 250 mL of 0.500M H₂SO₄ solution from 18M H₂SO₄?

For all problems show work in the format below. When you have the Know, Find, Formula information listed I will gladly help you.

- 1) 19.9 2) 106 3) 1.28 4) 1515 5) .288 6) .85 7) 54000 8) 7
9) .02 10) .75 11) .712 4.5 13) 8.3 14) .5 15) .875 and 29.125
16) 243.6 and 6.94

Know Find Formula Substitute & Solve

Name: _____ Date: _____ Block: _____

Making Dilutions Worksheet

Remember that you can change the concentration of a solution by adding more solvent. While you cannot increase the concentration of a solution in this manner, you can create a more dilute solution by increasing the amount of solvent. You can determine the amount of a solution needed to dilute by using the following: $M_1 \times V_1 = M_2 \times V_2$

Where M_1 = molarity and V_1 = volume. M_2 and V_2 are the initial solution molarity and volume, while M_1 and V_1 are the final solution molarity and volume.

Use the formula and information above to solve the following problems. Show your work and label each. The first three problems are questions regarding molarity and the others involve the dilution formula above.

- Determine the number of grams of NaHCO₃ that are in one liter of a 2.1 M solution.
- Determine the number of grams of NaNO₃ that are present in 500 mL of a 1.0 M solution.
- Determine the number of grams of CCl₄ in 450 mL of a 2.2 M solution.
- You have to make 500 mL of a 0.50 M BaCl₂. You have 2.0 M barium chloride solution available. Determine how to make the needed dilution.
- You need to make 200 mL of a 0.40 M solution of sodium chloride. The only available solution is 1.0 M. Determine how to make the needed dilution.

Dilution Name: _____

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